Name _____

Ions in Acids and Bases Answer Key

1. Dilution refers to the process of reducing the concentration of a solution by adding more solvent.

2. The chemical formula for nitric acid is HNO3.

3. The pH is 1. (Note: pH = 14 - pOH)

4. An amphiprotic substance is one that can act as both an acid and a base, depending on the reaction conditions.

5. In an acidic solution, the concentration of hydrogen ions is greater than the concentration of hydroxide ions.

6. A burette is used to deliver precise volumes of one solution into another during the titration process.

7. The pOH is 6. (Note: pOH = -log[OH-])

8. The pH is 8. (Note: pH = 14 - pOH)

9. Acid rain refers to rainwater or any precipitation with a pH lower than the neutral value of 7, primarily caused by air pollutants reacting with water in the atmosphere.

10. Buffering capacity refers to the ability of a solution to resist changes in pH when small amounts of acid or base are added to it.